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Strong versus Weak Acids
The following diagram illustrates the difference between strong and weak acids in water.

What
The diagram shows two beakers. The left beaker is labeled "Strong Acid" and contains a large number of dissociated ions (represented by small circles). The right beaker is labeled "Weak Acid" and contains a large number of undissociated acid molecules (represented by larger circles). A legend indicates that small circles represent "Water/Ions" and larger circles represent "Acid Molecules".

Model 1 - Acid Strength and Conductivity

The diagram illustrates the difference between strong and weak acids in water. Strong acids dissociate completely into ions, while weak acids do not. This is shown by the number of ions in the beakers. The strong acid beaker has many ions, while the weak acid beaker has few ions. The strong acid beaker also has a higher conductivity than the weak acid beaker.

1. Explain the difference between strong and weak acids in terms of their dissociation in water.

2. Explain the difference between strong and weak acids in terms of their conductivity.

3. Explain the difference between strong and weak acids in terms of their pH.

4. Explain the difference between strong and weak acids in terms of their reactivity.

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Answer Key Pogil Strong Versus Weak Acids